

Writing Compound Formulas Worksheet

Write the correct formula for each of the following ionic compound names.

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| 1. Copper (I) phosphate <u>Cu₃PO₄</u> | 16. Iron (III) sulfate <u>Fe₂(SO₄)₃</u> |
| 2. Magnesium nitrite <u>Mg(NO₂)₂</u> | 17. Silver dichromate <u>Ag₂Cr₂O₇</u> |
| 3. Cobalt(II) hydroxide <u>Co(OH)₂</u> | 18. Lead (II) nitride <u>Pb₃N₂</u> |
| 4. Iron (III) chromate <u>Fe₂(CrO₄)₃</u> | 19. Strontium hypochlorite <u>Sr(ClO)₂</u> |
| 5. Chromium (II) chloride <u>CrCl₂</u> | 20. Barium chlorate <u>Ba(ClO₃)₂</u> |
| 6. Zinc (II) fluoride <u>ZnF₂</u> | 21. Tin (IV) oxide <u>SnO₂</u> |
| 7. Mercury (II) acetate <u>Hg(C₂H₃O₂)₂</u> | 22. Nickel (II) cyanide <u>Ni(CN)₂</u> |
| 8. Manganese (II) iodide <u>MnI₂</u> | 23. Aluminum sulfide <u>Al₂S₃</u> |
| 9. Copper (II) nitrate <u>Cu(NO₃)₂</u> | 24. Lead (IV) chromate <u>Pb(CrO₄)₂</u> |
| 10. Aluminum hydroxide <u>Al(OH)₃</u> | 25. Potassium permanganate <u>KMnO₄</u> |
| 11. Iron (II) phosphate <u>Fe₃(PO₄)₂</u> | 26. Calcium dihydrogen phosphate <u>Ca(H₂PO₄)₂</u> |
| 12. Magnesium sulfite <u>MgSO₃</u> | - |
| 13. Tin (II) bromide <u>SnBr₂</u> | 27. Bariumhydrogensulfate <u>Ba(HSO₄)₂</u> |
| 14. Sodium bicarbonate <u>NaHCO₃</u> | 28. Lithium monohydrogen phosphate <u>Li₂HPO₄</u> |
| 15. Chromium(III)oxalate <u>Cr₂(C₂O₄)₃</u> | 29. Silver bisulfite <u>AgHSO₄</u> |
| | 30. Ammonium carbonate <u>(NH₄)₂CO₃</u> |

Write the Compound formulas for the covalent molecules below:

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| 31. Dinitrogen oxide <u>N₂O</u> | 35. Tetraphosphorus trisulfide <u>P₄S₃</u> |
| 32. Carbon tetraiodide <u>CI₄</u> | 36. Sulfur hexafluoride <u>SF₆</u> |
| 33. Dioxygenheptachloride <u>O₂Cl₇</u> | 37. Dinitrogen pentoxide <u>N₂O₅</u> |
| 34. Phosphorus tribromide <u>PBr₃</u> | 38. Tetrasulfur decanitride <u>S₄N₁₀</u> |

Write the compound name for the following covalent molecules.

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| 39. CS ₂ Carbon disulphide _____ | 43. N ₂ S ₅ <u>dinitrogen pentasulfide</u> |
| 40. PF ₅ <u>phosphorous pentafluoride</u> | 44. S ₂ F ₁₀ <u>disulfur decafluoride</u> |
| 41. CBr ₄ <u>Carbon tetra bromide</u> | 45. Cl ₂ O ₇ <u>dichlorine heptoxide</u> |
| 42. P ₄ S ₁₀ <u>tetraphosphorous decasulphide</u> | 46. I ₂ O ₅ <u>diiodine pentoxide</u> |